

The Chemistry of FIREWORKS

RED

STRONTIUM SALTS

ORANGE

CALCIUM SALTS

YELLOW

SODIUM COMPOUNDS

GREEN

BARIUM COMPOUNDS

BLUE

COPPER COMPOUNDS

PURPLE

STRONTIUM & COPPER COMPOUNDS

SILVER

BURNING ALUMINUM, AND TITANIUM

Luminescence:
light produced using energy sources - to produce luminescence, energy is absorbed by an electron of an atom or molecule, causing it to become excited, but unstable.

Incandescence:
Light produced from heat - heat causes a substance to become hot and glow, initially emitting infrared, then red, orange, yellow, and white light as it becomes increasingly hotter

What is needed?

COLOR PRODUCER

FUEL
allows firework to burn

OXIDISER
provides oxygen for the fuel combustion

BINDER
holds the mixture together